

CHEMISTRY

Welcome to Chemistry 101A  
Lecture 9:15 – 10:30 am  
PS H-150  
James P. Birk

Problems with registration, see:  
Shelley Morgan, H-233  
Dr. Ron Briggs, H-235

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CHEMISTRY

CHM-101

- Introductory course - no prerequisites
- Serves as a prerequisite for CHM 113/116 sequence or for CHM 231
- Satisfies lab science liberal arts requirement
- Lots of topics relating chemical concepts to our everyday lives

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Contact Info for CHM-101 A

- Instructor: J. Birk
- Lecture: 9:15-10:30 TTh
- Office: PS H-241
- Office Hours: 8:00-9:00 and 11:00-12:00 am TTh (Th 11:00 will have to be changed soon)
- Phone: (480) 965-3129
- E- Mail: JBirk@asu.edu
- Learning Resource Center (H-137) - TA office hours 8:30-9:30 M-Th, 8:30-4:30 F

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Handouts

- Syllabus
  - lecture schedule
  - lab schedule  
Note that labs start this week.
  - problem assignments
  - grading scheme
  - Corrections to textbook
- Class notes for Chapter 1
  - This and future chapter notes are available as pdf files on my web site (see syllabus)

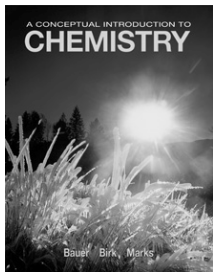
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Materials Needed

- From Bookstore
  - Text Book
  - Solutions manual (optional)
  - Lab Manual \*
- From SAACS / AXE or Bookstore
  - Goggles \*\*
  - Bound carbonless notebook for lab \*\*
- Binder for PowerPoint Notes
- Calculator

\* Need this week \*\* Need next week



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Lecture Schedule

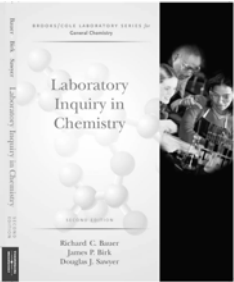
- Note dates for exams and final exam.
- Exams must be taken when scheduled.

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## Lab Schedule

- Lab starts this week.
- Lab experiment schedule is on syllabus



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## Group Work

- Lab/lecture group assignments to be made during the first lab meetings this week
- After groups are assigned, the group should sit together during lecture to facilitate discussions and group quizzes
- Starting next week, it will be assumed that groups are sitting together; group quizzes are possible any day from that time on.

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## Quizzes

- Discussion Quizzes (100 pts, 10 pts each)
  - Taken during discussion hour before lab, once a week.
  - Covers lecture material from the past week.
  - May be a question about the lab you're about to do.
  - We will post five end-of-chapter exercises every week. One of these will be the basis for a question on the next quiz.

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## Quizzes

- Lecture Quizzes (50 pts, 5 pts each)
  - Taken individually or in groups during lecture
  - Unannounced

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## Exams

- 4 Regular exams
  - 100 pts each
  - No Make-ups
- Final Exam
  - 200 pts
  - Resurrection Final

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## Course Point Distribution

● Hour Exams (4)	400 points
● Final Exam	200
● Labs	150
● Quizzes (12 or more)	100
● Group Quizzes (10 or more)	50
● TOTAL POINTS	900

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## Course Grade Distribution

Grade average >88%	A
>76%	B
>64%	C
>52%	D
52% or less	E

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## How to Use the Textbook

- Read the introduction to each chapter.
  - Shows how the chemistry in chapter is related to the real world.
- Read the chapter before class and again after class
- Do in-chapter and suggested end-of-chapter problems.
- Refer back to sections to help yourself with problems.

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## Ungraded Course Work

- To do well in this course, read and do textbook chapter problems on a regular basis.
- Textbook Suggested Problems
  - Minimum effort; do more if needed
  - In syllabus

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## Learning Resource Center (LRC)

- The Chemistry LRC (H-137) will open in the next couple of days
  - TA office hours
  - Table to study with other students
  - Computers for chemistry work
  - Other resources such as old textbooks
- The University LRC will sponsor Directed-Study Sessions led by student Team Leaders. More info will be provided during discussion sessions this week. Also, see p. 4 of the syllabus.

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## Web site

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## General Chemistry

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James P. Birk  
Office: PS 10-241  
Phone: (480) 960-3329

### CHM-101 Introductory Chemistry

Get Instructor's  
Materials

Lecture Notes	A Conceptual Introduction to Chemistry	Review, Exts. & Maps
Introduction		
Syllabus		
Chapter 1	2-42	Matter and Energy
Chapter 2	40-72	Atoms, Ions, and the Periodic Table
Chapter 3	70-108	Chemical Compounds
Chapter 4	115-148	Chemical Composition
Chapter 5	152-184	Chemical Reactions and Equations
Chapter 6	194-227	Quantities in Chemical Reactions
Chapter 7	230-274	Electron Structure of the Atom
Chapter 8	276-310	Chemical Bonding
Chapter 9	314-362	The Gaseous State
Chapter 10	362-399	The Liquid and Solid States
Chapter 11	406-440	Solutions
Chapter 12	446-480	Reaction Rates and Chemical Equilibrium
Chapter 13	489-518	Acids and Bases
Return		Return to course selection

Updated on: 7/26/2006  
For comments or feedback contact: James P. Birk.  
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