

CHEMISTRY

Welcome to Chemistry 101A  
Lecture 9:15 – 10:30 am  
PS H-150  
James P. Birk

Problems with registration, see:  
Shelley Morgan, H-233  
Dr. Ron Briggs, H-235

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CHM-101

- Introductory course - no prerequisites
- Serves as a prerequisite for CHM 113/116 sequence or for CHM 231
- Satisfies lab science liberal arts requirement
- Lots of topics relating chemical concepts to our everyday lives

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Contact Info for CHM-101 A

- Instructor: J. Birk
- Lecture: 9:15-10:30 TTh
- Office: PS H-241
- Office Hours: 8:00-9:00 and 11:00-12:00 am TTh (Th 11:00 will have to be changed soon)
- Phone: (480) 965-3129
- E- Mail: JBirk@asu.edu
- Learning Resource Center (H-137) - TA office hours 8:30-9:30 M-Th, 8:30-4:30 F

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Handouts

- Syllabus
  - lecture schedule
  - lab schedule  
Note that labs start this week.
  - problem assignments
  - grading scheme
  - Corrections to textbook
- Class notes for Chapter 1
  - This and future chapter notes are available as pdf files on my web site (see syllabus)

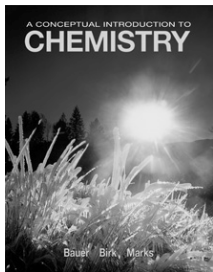
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Materials Needed

- From Bookstore
  - Text Book
  - Solutions manual (optional)
  - Lab Manual \*
- From SAACS / AXE or Bookstore
  - Goggles \*\*
  - Bound carbonless notebook for lab \*\*
- Binder for PowerPoint Notes
- Calculator

\* Need this week \*\* Need next week



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Lecture Schedule

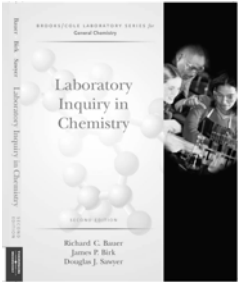
- Note dates for exams and final exam.
- Exams must be taken when scheduled.

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## Lab Schedule

- Lab starts this week.
- Lab experiment schedule is on syllabus



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## Group Work

- Lab/lecture group assignments to be made during the first lab meetings this week
- After groups are assigned, the group should sit together during lecture to facilitate discussions and group quizzes
- Starting next week, it will be assumed that groups are sitting together; group quizzes are possible any day from that time on.

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## Quizzes

- Discussion Quizzes (100 pts, 10 pts each)
  - Taken during discussion hour before lab, once a week.
  - Covers lecture material from the past week.
  - May be a question about the lab you're about to do.
  - We will post five end-of-chapter exercises every week. One of these will be the basis for a question on the next quiz.

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## Quizzes

- Lecture Quizzes (50 pts, 5 pts each)
  - Taken individually or in groups during lecture
  - Unannounced

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## Exams

- 4 Regular exams
  - 100 pts each
  - No Make-ups
- Final Exam
  - 200 pts
  - Resurrection Final

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## Course Point Distribution

|                              |            |
|------------------------------|------------|
| ● Hour Exams (4)             | 400 points |
| ● Final Exam                 | 200        |
| ● Labs                       | 150        |
| ● Quizzes (12 or more)       | 100        |
| ● Group Quizzes (10 or more) | 50         |
| ● TOTAL POINTS               | 900        |

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## Course Grade Distribution

|                    |   |
|--------------------|---|
| Grade average >88% | A |
| >76%               | B |
| >64%               | C |
| >52%               | D |
| 52% or less        | E |

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## How to Use the Textbook

- Read the introduction to each chapter.
  - Shows how the chemistry in chapter is related to the real world.
- Read the chapter before class and again after class
- Do in-chapter and suggested end-of-chapter problems.
- Refer back to sections to help yourself with problems.

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## Ungraded Course Work

- To do well in this course, read and do textbook chapter problems on a regular basis.
- Textbook Suggested Problems
  - Minimum effort; do more if needed
  - In syllabus

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## Learning Resource Center (LRC)

- The Chemistry LRC (H-137) will open in the next couple of days
  - TA office hours
  - Table to study with other students
  - Computers for chemistry work
  - Other resources such as old textbooks
- The University LRC will sponsor Directed-Study Sessions led by student Team Leaders. More info will be provided during discussion sessions this week. Also, see p. 4 of the syllabus.

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## Web site

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## General Chemistry

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James P. Birk  
Office: PS 10-241  
Phone: (480) 960-3329

### CHM-101 Introductory Chemistry

Get Instructor's  
Materials

| Lecture Notes | A Conceptual<br>Introduction to Chemistry | Review, Exts. & Maps                    |
|---------------|---|---|
| Introduction  |   |   |
| Syllabus      |   |   |
| Chapter 1     | 2-42                                      | Matter and Energy                       |
| Chapter 2     | 40-72                                     | Atoms, Ions, and the Periodic Table     |
| Chapter 3     | 70-108                                    | Chemical Compounds                      |
| Chapter 4     | 115-148                                   | Chemical Composition                    |
| Chapter 5     | 152-184                                   | Chemical Reactions and Equations        |
| Chapter 6     | 194-227                                   | Quantities in Chemical Reactions        |
| Chapter 7     | 230-274                                   | Electron Structure of the Atom          |
| Chapter 8     | 276-310                                   | Chemical Bonding                        |
| Chapter 9     | 314-362                                   | The Gaseous State                       |
| Chapter 10    | 362-399                                   | The Liquid and Solid States             |
| Chapter 11    | 406-440                                   | Solutions                               |
| Chapter 12    | 446-480                                   | Reaction Rates and Chemical Equilibrium |
| Chapter 13    | 489-518                                   | Acids and Bases                         |
| Return        |   | Return to course selection              |

Updated on: 7/26/2006  
For comments or feedback contact: James P. Birk.  
#Birk.Address: jpbirk@asu.edu

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